**Name of the Student**:_____
**Max. Marks : 16 Marks** **Time : 16 Minutes**

Mark Schemes

Q1.

(a) Elements in the p block have their outer electron(s) in
p orbital(s) or levels or sub-shells **(1)**
example of element **(1)**
correct electronic configuration **(1)**

**3**

(b) Pattern in the change in the properties of a row of elements **(1)**
*OR Trend in the properties of elements across a period*

Repeated in the next row **(1)**
*OR element underneath (or in same group) has similar properties*

**atomic radius**
decreases across the row **(1)**
*CE if trend is wrong*

number of protons increases **(1)** (or nuclear charge increases)
more attraction for electrons in the same shell **(1)**

**electronegativity**
increases across the row **(1)**
number of protons increases **(1)** (or nuclear charge)
atomic radius decreases **(1)** (or shielding remains the same or electrons in the same shell) more attraction for bonding or shared electrons **(1)**

**conductivity**
decreases row **(1)**
*OR significant drop from Al to Si*

Na–Al metals **(1)**
*OR metallic bonding or description of metallic bonding*

Two of Si - Ar non metals **(1)**
*OR molecular or covalent*

EITHER electrons free to move (or delocalised) in metals
OR electrons unable to move in non-metals **(1)**

**13**

**[16]**

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