Practice Question Set For A-Level
Subject : Chemistry
Paper-1 Topic: Inorganic 2

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Name of the Student:_____
Max. Marks : 25 Marks **Time : 25 Minutes**

Mark Schemes

Q1.

(a)

This question is marked using levels of response. Refer to the Mark Scheme Instructions for Examiners for guidance on how to mark this question.	
Level 3 **5-6 marks**	All stages are covered and the description of each stage is generally correct and virtually complete. Answer is communicated coherently and shows a logical progression from stage 1 to stage 2 and stage 3.
Level 2 **3-4 marks**	All stages are covered but the description of each stage may be incomplete or may contain inaccuracies **OR** two stages are covered and the explanations are generally correct and virtually complete. Answer is mainly coherent and shows progression from stage 1 to stage 2 and/or stage 3.
Level 1 **1-2 marks**	Two stages are covered but the description of each stage may be incomplete or may contain inaccuracies, **OR** only one stage is covered but the explanation is generally correct and virtually complete. Answer includes isolated statements and these are presented in a logical order.
Level 0	**0 marks** Insufficient correct chemistry to gain a mark.

Stage 1
1a Heterogeneous means in a different phase/state from reactants
1b Catalyst speeds up reaction and is left unchanged **OR** lowers the activation energy for the reaction

Stage 2
2a Hydrogen and nitrogen/reactants adsorb onto the surface/ active sites of the iron
2b Bonds weaken/reaction takes place
2c Products desorb/leave from the surface (of the iron)

Stage 3
3a Large surface area (of iron) by using powder or small pellets or support medium/mesh
3b Catalyst poisoned / sulfur poisons or binds to the catalyst
3c Active sites blocked

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Ignore references to temperature and pressure

6

(b) Two negative ions repel

1

So activation energy is high

1

$2\ Fe^{2+} + S_2O_8^{2-} \rightarrow 2\ SO_4^{2-} + 2\ Fe^{3+}$

1

$2\ Fe^{3+} + 2\ I^- \rightarrow 2\ Fe^{2+} + I_2$

Ignore any state symbols given
Allow multiples for both equations
Allow equations in either order

1

(c) (Zn ions) have only one oxidation state
Or
Zn^{2+} is the only ion

Allow doesn't have variable oxidation state
Allow cannot be oxidised to Zn^{3+}
Ignore has a full d shell

1

(d) M1 Amount of Fe = 0.998 ÷ 55.8 = 0.0179 mol

1

M2 Amount of HCl = 0.0300 mol

1

M3 HCl is the limiting reagent

1

M4 Amount of H_2 produced = 0.0150 mol

M4 = M2÷2

1

M5 T = 303 K P = 100 000 Pa

1

M6 $V \left(= \frac{0.0150 \times 8.31 \times 303}{100\ 000}\right) = 3.78 \times 10^{-4}\ (m^3)$

M6 $V \left(= \frac{M4 \times 8.31 \times 303}{100\ 000}\right) (m^3)$

1

(e) $FeCO_3$ or iron(II) carbonate

1

Green

Allow white

1

(f) $Fe(H_2O)_3(OH)_3$

Ignore square brackets if added

1

brown

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1

$2\ [Fe(H_2O)_6]^{3+} + 3\ CO_3^{2-} \rightarrow 2\ Fe(H_2O)_3(OH)_3 + 3\ H_2O + 3\ CO_2$

Accept multiples

1

(g) M1 Fe^{3+} is smaller (than Fe^{2+}) **OR** Fe^{3+} has a greater charge **OR** Fe^{3+} has a greater charge density **OR** Fe^{3+} has a greater charge to size ratio

Penalise $Fe(H_2O)_6^{3+}$ ions once in M1 or M2

1

M2 Fe^{3+} ions are more polarising **OR** Fe^{3+} ions polarise water molecules more

1

M3 So more O-H bonds (in the water ligands) break **OR** more H^+ ions released **OR** weaken O-H bonds in ligands more (in the Fe^{3+} solution)

Do not allow Fe^{3+} releases $3H^+$ ions

1

[25]

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